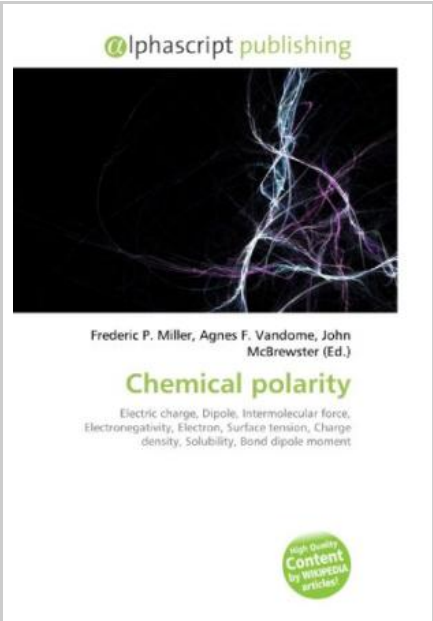

Chemical polarity

By Frederic P. Miller

Alphascript Publishing Dez 2009, 2009. Taschenbuch. Book Condition: Neu. 220x150x7 mm. Neuware - High Quality Content by WIKIPEDIA articles! In chemistry, polarity refers to a separation of electric charge leading to a molecule having an electric dipole. Polar molecules can bond together due to dipole. Dipole intermolecular forces between one molecule (or part of a large molecule) with asymmetrical charge distribution and another molecule also with asymmetrical charge distribution. Molecular polarity is dependent on the difference in electronegativity between atoms in a compound and the asymmetry of the compound's structure. For example, a molecule of water is polar because of the unequal sharing of its electrons in a 'bent' structure, whereas methane is considered non-polar because the carbon shares the electrons with the hydrogen atoms uniformly. Polarity underlies a number of physical properties including surface tension, solubility, and melting and boiling-points. 120 pp. Englisch.

Reviews

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